

1. The ice of an outdoor skating rink is at the temperature of  $-2.0^\circ\text{C}$ . Calculate the minimum pressure (applied for example by a skate) necessary to melt the ice. (20%)  
Data: At  $0^\circ\text{C}$ , the specific volume of water is  $1.000\text{ cm}^3/\text{g}$  and that of ice is  $1.090\text{ cm}^3/\text{g}$ ; the heat of fusion is  $79.7\text{ cal/g}$ .  
Unit conversions:  $1\text{ atm} = 101325\text{ N/m}^2$ ,  $1\text{ cal} = 4.18\text{ J}$

2. Show that the slope of  $\beta$  vs.  $T$  curve is horizontal at  $T = 0\text{ K}$ , i.e.  $\left(\frac{\partial\beta}{\partial T}\right)_P = 0$  at  $T = 0\text{ K}$ . (20%)

Where  $\beta =$  the coefficient of isothermal compressibility  $= -\frac{1}{V}\left(\frac{\partial V}{\partial P}\right)_T$ ,  $T =$  temperature,  $V =$  volume,  
 $P =$  pressure

3. Al-Zn alloys exhibit the following relation at  $477^\circ\text{C}$

$$RT \ln \gamma_{\text{Zn}} = 1750(1 - X_{\text{Zn}})^2$$

Where  $R = 1.987\text{ cal/mole K}$ , and  $T$  is expressed in K.  $\gamma_{\text{Zn}}$  and  $X_{\text{Zn}}$  are, respectively, the activity coefficient and the molar fraction of Zn in the alloy. Calculate the activity of aluminum at  $477^\circ\text{C}$  in an Al-Zn alloy containing 40 atom% zinc. (20%)

4. Nickel and element A are related by a phase diagram of the type shown in Figure 1. At  $727^\circ\text{C}$ , element A will dissolve 4 atom% nickel and nickel will dissolve 6 atom% A. Calculate the oxygen pressure at which nickel dissolved in A to the extent of 1 atom% will just begin to oxidize to NiO. Element A is more stable than nickel and will oxidize only at higher oxygen pressures. (20%)  
Given: the standard free energy for the reaction

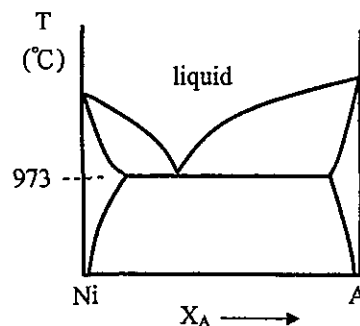
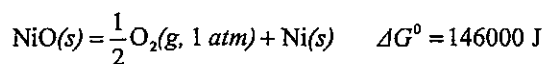


Figure 1

5. The molar free energy,  $G$ , of a regular solution composed of elements A and B can be expressed as  
$$G = X_A G_A + X_B G_B + \Omega X_A X_B + RT(X_A \ln X_A + X_B \ln X_B)$$
  
Where  $G_A$  and  $G_B$  are the molar free energy of pure A and B, respectively.  $X_A$  and  $X_B$  are the molar fraction of A and B in the solution.  $\Omega$  is related to the difference between A-B bonds and (A-A bonds + B-B bonds).

- (1) Show that  $\frac{d^2G}{dX^2} = \frac{RT}{X_A X_B} - 2\Omega$ , where  $X$  is  $X_A$  or  $X_B$ . (10%)

- (2) Based on the equation shown in question 5.(1), what are the conditions under which the free energy curve ( $G$  vs.  $X_B$  plot) of a regular solution exhibits a maximum? (10%)

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